**Chemistry Stoichiometry (Chapter 11 Outline)**

***AS YOU READ, RESPOND TO THE FOLLOWING:***

*Vocabulary*: **Define in your own words**

stoichiometry-

mole ratio-

limiting reactant-

excess reactant-

theoretical yield-

actual yield-

percent yield-

Respond to the prompts below:

How are mole ratios written from a balanced chemical equation? Explain. (2 pts)

What is the sequence of steps that should be used to solve stoichiometric problems? Explain how these steps are applied to solving stoichiometric problems. (3 pts)

In a chemical reaction, which reactant is the limiting reactant? Explain how you determine the limiting reactant in a chemical reaction. (2 pts)

How do you calculate the mass of a product when the amounts of more than one reactant are given? Explain. (2 pts)

How do you calculate the percent yield for a chemical reaction? Explain. (2 pts)

***ABOUT THE READING:***

Write three things that you learned about Stoichiometry*:*

***Make sure to write a full sentence.***

*Example: I learned that temperature is a measure of an object's kinetic energy.*

*1.*

*2.*

*3.*

***Assigned work****: Chapter 11 assessment, problems #48-50, 64-67, 77-79, 81, 90-93*